

1 **Technical Note: Stability of tris pH buffer in artificial seawater**
2 **stored in bags**

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9

10 **Abstract**

11 Equimolar tris (2-amino-2-hydroxymethyl-propane-1,3-diol) buffer in artificial seawater is a well characterized
12 and commonly used standard for oceanographic pH measurements. We evaluated the stability of tris pH when stored
13 in purportedly gas impermeable bags across a variety of experimental conditions, including bag type, and storage in
14 air vs. seawater over 300 days. Bench-top spectrophotometric pH analysis revealed that the pH of tris stored in bags
15 decreased at a rate of $0.0058 \pm 0.0011 \text{ yr}^{-1}$ (mean slope \pm 95% confidence interval of slope). The upper and lower
16 bounds of expected pH change at $t = 365$ days, calculated using the averages and confidence intervals of slope and
17 intercept of measured pH change vs. time data, were -0.0042 and -0.0076 from initial pH. Analyses of total dissolved
18 inorganic carbon confirmed that a combination of CO₂ infiltration and/or microbial respiration led to the observed
19 decrease in pH. Eliminating the change in pH of bagged tris remains a goal, yet the rate of pH change is lower than
20 many processes of interest and demonstrates the potential of bagged tris for sensor calibration and validation of
21 autonomous in situ pH measurements.

22 **1. Introduction**

23 Ocean pH is a key measurement used for tracking biogeochemical processes such as photosynthesis,
24 respiration, and calcification (Takeshita et al., 2016); and represents perhaps the most recognized variable associated
25 with ocean acidification (OA), the decrease in ocean pH due to the uptake of anthropogenic carbon dioxide (Doney et
26 al., 2009). OA progresses with a global average pH decline of 0.002 per year in the surface open ocean (Bates et al.,
27 2014), and the accumulated and projected near-term effects of OA have been shown to have deleterious effects on
28 many calcifying organisms (Cooley and Doney, 2009). Beyond the narrow scope of calcifiers, organismal response is
29 complex, exhibiting varied responses across processes such as reproduction, growth rate, and sensory perception.
30 Organismal responses are further complicated by their impact on ecosystem level dynamics, such as altering
31 competition and predator-prey relationships (Doney et al., 2020). Furthermore, pH effects are often exacerbated by
32 concomitant stressors, such as decreased dissolved oxygen or increased temperature. Ultimately, OA will affect
33 humans through impacts on fisheries, aquaculture, and shoreline protection (Branch et al., 2013; Doney et al., 2020).

34 The quality of pH measurement required to observe various phenomena is often broken into “climate” and
35 “weather” levels of uncertainty (Newton et al., 2015), or 0.003 and 0.02, respectively. Discrete sampling has been
36 shown to be capable of meeting the climate level of uncertainty when best practices are followed, yet many labs do
37 not consistently meet this standard (Bockmon and Dickson, 2015). Furthermore, while discrete, bench-top
38 methodologies can be the most accurate, the ocean’s vast size limits the oceanographic community’s ability to make
39 ship-based discrete pH measurements to decadal reoccupations of a few major sections per ocean basin (Sloyan et al.,
40 2019). The sparsity of ship-board measurements hinders our ability to assess sub-decadal processes, such as seasonal
41 cycles or bloom events, over much of the ocean (Karl, 2010), and highlights the need for autonomous, high-frequency
42 pH measurements. Technological advancements have led to more routine autonomous pH measurements over the past
43 decade, providing opportunities to fill some gaps in time and space in discrete sampling programs (e.g. Byrne, 2014;
44 Martz et al., 2015; Lai et al., 2018; Wang et al., 2019; Tilbrook et al., 2019). Globally, pH sensors now operate on

45 hundreds of autonomous platforms including moorings and profiling floats, delivering unique datasets in the form of
46 Eulerian and depth resolved Lagrangian time series (Johnson et al., 2017; Bushinsky et al., 2019; Sutton et al., 2019).
47 While sensors increase data coverage, many sensor-based pH measurements, particularly on moored systems, continue
48 to fall short of both climate and weather levels of uncertainty, as highlighted in the intercomparison tests carried out
49 by the Alliance for Coastal Technologies (ACT, 2012) and by the Wendy Schmidt Ocean Health XPRIZE (Okazaki
50 et al., 2017).

51 Independent validation is typically required for autonomous sensors to meet both weather and climate levels
52 of uncertainty. For example, autonomous underway $p\text{CO}_2$ systems (Pierrot et al., 2009), moorings (Bushinsky et al.,
53 2019), and autonomous surface vehicles (Chavez et al., 2017; Sabine et al., 2020) are able to provide climate quality
54 observations with an uncertainty of $\pm 2 \mu\text{atm}$ because traceable standard gases are frequently measured in situ. For pH
55 measurements on profiling floats (Johnson et al., 2016), sensor performance is validated by comparing to a deep
56 reference pH field that is calculated using empirical algorithms (Williams et al., 2016; Bittig et al., 2018; Carter et al.,
57 2018). This approach has demonstrated the ability to obtain high quality pH measurements from a network of profiling
58 floats (Johnson et al., 2017) but requires measurements in the deep ocean where pH is comparatively stable. It is
59 atypical for other pH sensors, including coastal moored sensors, to have an automated or remote validation. Therefore,
60 on such deployments, validation has largely relied on discrete samples taken alongside the sensor (Bresnahan et al.,
61 2014; McLaughlin et al., 2017; Takeshita et al., 2018), which presents unique challenges; primarily that spatiotemporal
62 discrepancy can lead to errors of > 0.1 , especially in highly dynamic systems (Bresnahan et al., 2014).

63 Similar to the method in use by $p\text{CO}_2$ systems, one approach to validate in situ pH sensors is by measuring a
64 reference material or pH standard, repeatedly during a sensor deployment. The most commonly used standard for
65 oceanographic pH measurement is equimolar tris (2-amino-2-hydroxymethyl-propane-1,3-diol) buffer in artificial
66 seawater (ASW), hereafter referred to as tris or tris-ASW (DeValls and Dickson, 1998; Papadimitriou et al., 2016).
67 The pH of tris has been characterized over a range of temperature, salinity, and pressure (DeValls and Dickson, 1998;
68 Rodriguez et al., 2015; Takeshita et al., 2017; Müller et al., 2018), allowing for accurate calculation of tris pH across
69 a wide range of marine conditions. Furthermore, when stored in borosilicate bottles and under ideal conditions, these
70 buffers have been shown to be stable to better than 0.0005 over a year (Dickson, 1993; Nemzer and Dickson, 2005),
71 making tris a good candidate for in situ validation of long-term deployments of autonomous pH sensors. To be utilized
72 for in situ applications, the reference solution must be stored in bags (as in, Hales et al., 2005; Seidel et al., 2008;
73 Sayles and Eck, 2009; Spaulding et al., 2014; Wang et al., 2015; Lai et al., 2018). Recently, in situ sensor validation
74 using bagged tris was demonstrated by Lai et al. (2018) during a 150-day deployment of an autonomous pH sensor,
75 where the tris standard was measured in situ every 5 days. However, the stability of tris when stored in bags has not
76 been quantified systematically using spectrophotometric bench-top pH measurement techniques recommended as best
77 practices (Dickson et al., 2007).

78 In this work we quantified the stability of tris stored in bags for 300 days. Tris from four separately prepared
79 batches was stored in two bag types either in a lab or submerged in seawater. In addition, one batch was stored in
80 borosilicate bottles in the lab as a control. Spectrophotometric pH measurements were made approximately every two

81 months on each bag of tris. Throughout the experiment, Certified Reference Materials (CRMs) for oceanic CO₂
82 measurements (Dickson, 2001) were used to assess the stability of the spectrophotometric pH system.

83 2. Methods

84 Two bag types were tested for storing tris (Figure 1). Bag type 1 was custom made based on a design used in
85 the “Burke-o-Lator” system (Hales et al., 2005; Bandstra et al., 2006), made from PAKDRY 7500 barrier film
86 (IMPAK P75C0919). The barrier film is made of layers of polyester and nylon with a sealant layer of metallocene
87 polyethylene. Two 23 x 48 cm (9” x 19”) sheets were heat sealed on three sides, forming a pocket, and a 1.9 cm (3/4”)
88 diameter hole was cut into one of the pocket walls for the bulkhead fitting and bulkhead nut (McMaster-Carr
89 8674T55). The bulkhead was sealed into the wall with a silicone gasket (McMaster-Carr 9010K13), washer
90 (McMaster-Carr 95649A256), and coated with silicone sealant (McMaster-Carr 74955A53). A “push-to-connect” ball
91 valve fitting (McMaster-Carr 4379K41), was attached to the bulkhead. The modified pocket was rinsed, dried, and
92 heat sealed along the final edge to create a ~4 L bag. Bags were left to dry for at least 24 hours before filling. Bag type
93 2 was a commercially available 3 L Cali-5-Bond bag purchased from Calibrated Instruments and used without
94 modification. It is a multi-layer bag made of plastic, aluminium foil (to prevent liquid and gas permeation), a layer of
95 inert high density polyethylene (to form a non-reactive inner wall) and, a polycarbonate Stopcock Luer valve.

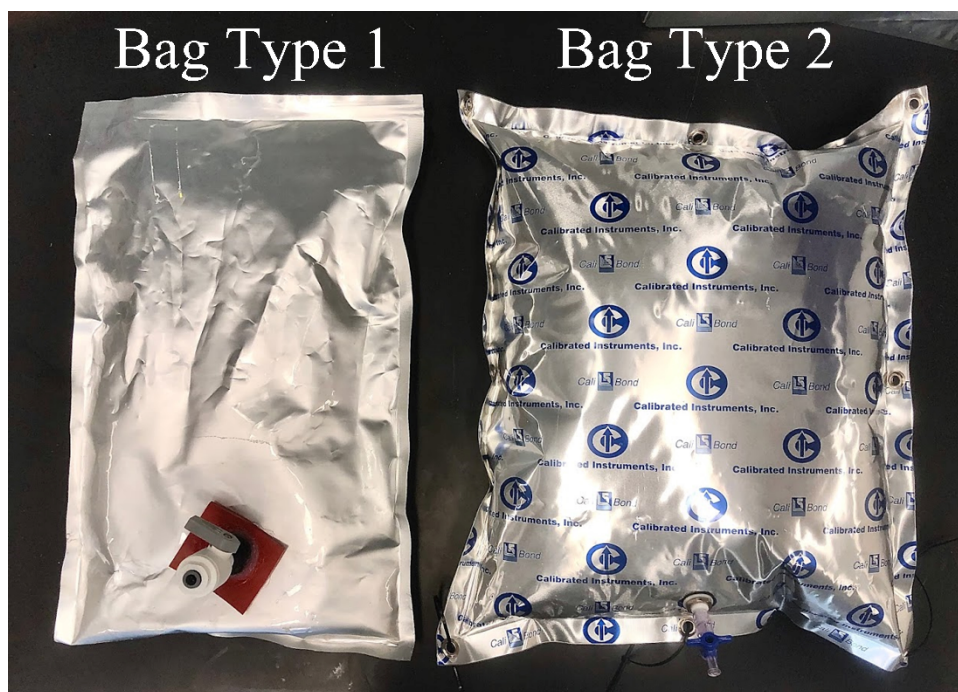


Figure 1: A picture of bag type 1 and 2 used to store tris in this study.

96 In this experiment, four batches of tris were prepared following the procedure in DelValls and Dickson
97 (1998), using off-the-shelf reagents with no additional standardization or purification (e.g. recrystallization of salts).
98 The focus of this paper is stability of bagged tris over time and does not prioritize obtaining highly accurate equimolar
99 tris (as would be necessary for characterization of thermodynamic constants, for example). The calculated pH of tris

100 in this study was 8.2652 at 20°C, based on quantity of reagents used. This is 0.0135 higher than the pH of equimolar
 101 tris, 8.2517 at 20°C (DeValls and Dickson, 1998). The pH discrepancy was due to a unit error in the measurement of
 102 HCl (our preparation used mol/L rather than the prescribed mol/kg-sol). This unit error resulted in a tris:trisH⁺ of
 103 1:0.97 that slightly differs from the 1:1 of truly equimolar tris. As this ratio is nearly equimolar, the term “equimolar”
 104 will continue to be used throughout this study. The details of the specific reagents used to prepare the tris solution can
 105 be found in Table A1.

106 Three stability tests were initiated at different times over the course of 18 months. The initiation of a given
 107 test is defined as the date of preparation of the tris used in that test. A summary of the differences between these tests
 108 is shown in Table 1 and described here. Each bag has a unique identifier in the format of “Batch #, Bag #, Lab or
 109 Tank.” If this identifier is duplicative, the bags are differentiated with letters A to D. Each bag was rinsed before
 110 filling: 3 times with deionized water (DI), 5 times with ultrapure water (> 18 MΩ resistivity) and at least 3 times with
 111 200 mL of tris. Tris bags were stored on a lab bench or in a 5,000 L test tank filled with ozone-sterilized, filtered
 112 seawater. Bag type 2 experienced delamination of exterior layers when stored in seawater during test 2 and was not
 113 used in further testing. Tris from batch 4 was also stored in borosilicate bottles following the procedure in Nemzer
 114 and Dickson (2005). In addition to pH measurements, dissolved inorganic carbon (C_T) was measured on both bagged
 115 and bottled tris during test 3 to see if changes in pH were due to increased CO₂. C_T samples were measured using a
 116 custom-built system based on an infrared (IR) analyser (LI-COR 7000) similar to systems used by O’Sullivan and
 117 Millero (1998) and Friederich et al. (2002). This IR measurement system is capable of measuring relatively low C_T
 118 without requiring method adjustment and has been used to make near zero C_T measurements (Paulsen and Dickson,
 119 unpublished data). C_T measurements were made on CRMs (Batch 179 & 183). The precision of the C_T measurements
 120 was $\pm 1.4 \mu\text{mol/kg}$ (pooled standard deviation, $n_{\text{samples}}=15$, $n_{\text{measurements}}=44$).

121 **Table 1: Tris preparation and storage.**

	Bag Type	Tris Batch	Date Made	Storage Location	Rinse Procedure	C_T Measured
Test 1	1 & 2	1 & 2	13 Dec 2017	Lab & Tank	3x DI, 5x ultrapure, 3x tris	No
Test 2	1 & 2	3	13 April 2018	Lab & Tank	3x DI, 5x ultrapure, 3x tris	No
Test 3	1 & bottle	4	26 February 2019	Lab	3x DI, 5x ultrapure, $\geq 6x$ tris	Yes

122
 123 Tris pH was measured every 55 ± 20 days (mean \pm standard deviation of measurement interval) throughout
 124 the experiment. The pH of tris was measured in triplicate at each time point with spectrophotometry using m-cresol
 125 purple as the indicator dye using the system described in Carter et al. (2013). Absorbance measurements were made
 126 in a 10-cm jacketed cell, and the temperature was measured directly adjacent to the cell outflow using a NIST-traceable
 127 thermometer (± 0.1 °C, QTI DTU6028P-001-SC). Blank and sample were held for 3 minutes in the jacketed flow cell
 128 prior to absorbance measurements.

129 On average, temperature was stable to within a 0.02 °C range over the course of the day; the mean temperature
 130 throughout the experiment was 20.09 ± 0.23 °C (1 σ), although temperature was 0.6 °C higher than the average on
 131 one measurement day. Spectrophotometric pH measurements are reported at 20 °C by adjusting the measured pH
 132 value at the measured cell temperature T_C (pH_{spec,T_C}) to 20 °C ($pH_{spec,20^\circ C}$) using the known temperature dependence
 133 of tris (pH_{tris}) as follows:

$$pH_{spec,20^\circ C} = pH_{spec,T_C} - (pH_{tris,T_C} - pH_{tris,20^\circ C}) \quad (1)$$

134 pH_{tris,T_C} and $pH_{tris,20^\circ C}$ are the theoretical pH of tris (at the measured temperature and 20 °C respectively) and were
 135 calculated using Eq. (18) in DelValls and Dickson (1998). This adjustment assumes that any potential difference in
 136 $\partial pH/\partial T$ between that corresponding to equimolar tris and that corresponding to our 1:0.97 tris:trisH⁺ ratio has a
 137 negligible effect over the small temperature range observed.

138 To account for pH-dependent errors from impurities in unpurified mCP, a pH-dependent correction factor
 139 was determined based on the protocol outlined in Takeshita et al. (*in review*). Briefly, pH of natural seawater with
 140 different ratios of added tris:trisH⁺ was measured subsequently using impure dye (pH_{impure} ; from Aldrich, lot
 141 MKBH6858V) and purified dye (pH_{pure} ; from Robert Byrne’s laboratory, University of South Florida (Liu et al.,
 142 2011)) over a range of pH between 7.4 to 8.2 at approximately 0.2 intervals. Varying ratios of tris:trisH⁺ were used to
 143 obtain different solution pH, and to buffer any changes in pH during the experiment, which negates the need for dye
 144 perturbation corrections in this characterization. Triplicate measurements were made at each pH. A second order pH-
 145 dependent error was observed as previously described, following the equation ($R^2 = 0.975$, RMSE = 0.000434):

$$pH_{pure} = -0.0047777 \times pH_{impure}^2 + 1.0668875 \times pH_{impure} - 0.2359740 \quad (2)$$

146 All subsequent pH_{spec} measurements in this study were conducted with impure dye and are reported with this dye
 147 impurity correction (Eq. 2) applied. The correction adjusted the reported pH by 0.0093 ± 0.0002 (mean \pm standard
 148 deviation, $n = 126$). No dye perturbation correction was used (a correction for a change in pH caused by the addition
 149 of the dye). As the high buffering capacity of tris, in combination with a dye adjusted to a pH similar to that of tris,
 150 results in a negligible change in measured pH.

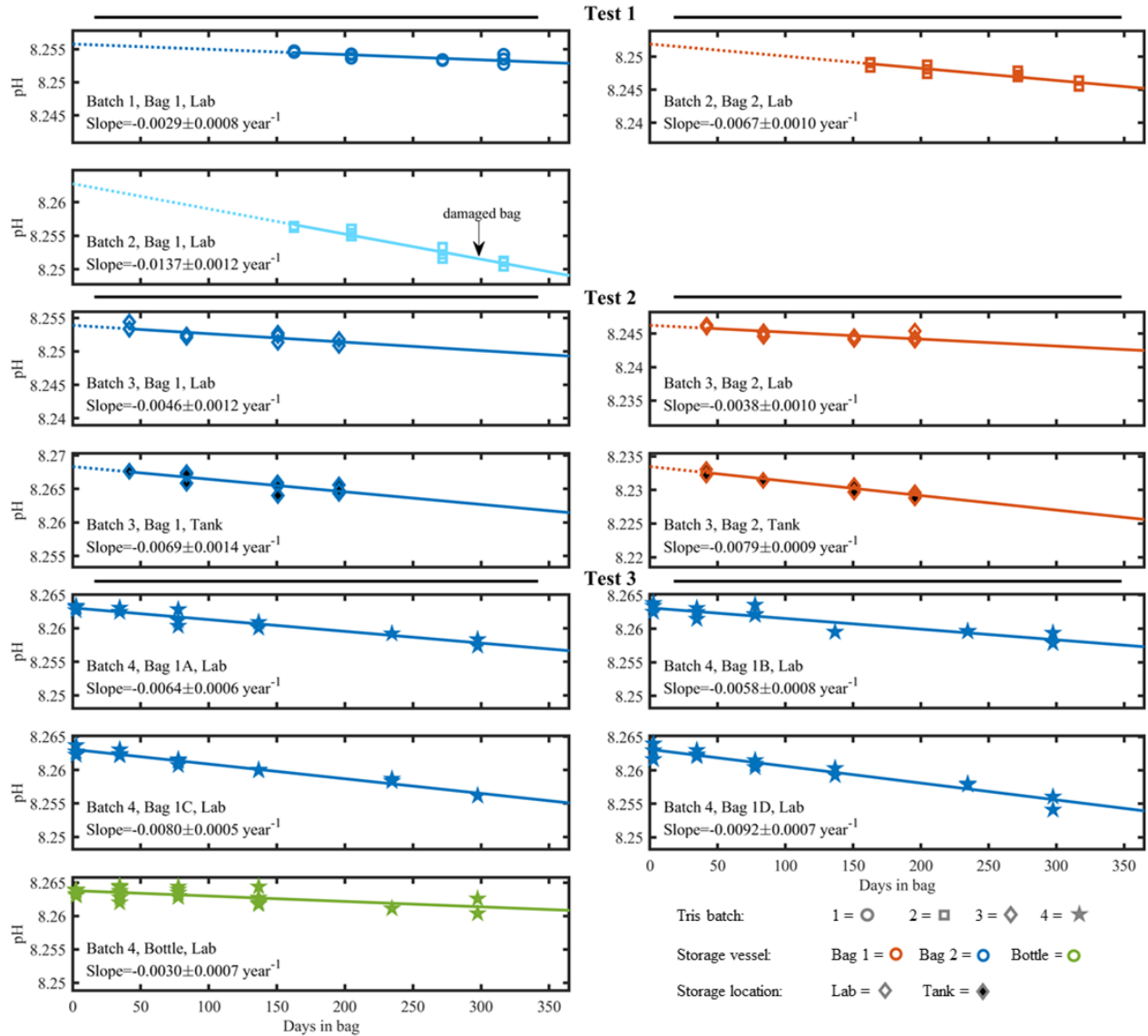
151 Measurements of tris batches 1 and 2 made in the first 150 days have been removed from the data set due to
 152 procedural changes made to the spectrophotometric pH system to correct for problems with temperature equilibration.
 153 Outliers were removed from the spectrophotometric pH measurements if the absorbance at 760 nm was above 0.005
 154 or below -0.002 (indicative of a measurement problem, such as a bubble or lamp drift), resulting in the removal of 2
 155 out of 163 measurements. Additionally, outliers were removed from the data set if they were greater than three
 156 standard deviations from the mean of a measurement triplicate, where standard deviation is calculated as using all sets
 157 of triplicates (1 standard deviation = 0.0004, $n = 55$), resulting in the removal of 2 of 161 remaining measurements.
 158 The remaining 159 measurements were used for the analysis presented here. An analysis of variation, or ANOVA,
 159 was used to detect the dependence of the results on tris batch, bag/bottle type and storage location. Analysis was
 160 performed using MATLAB R2020a and the standard function “*anovan()*.” Throughout the experiment, CRMs
 161 (procured from A. Dickson, Scripps Institution of Oceanography) for seawater C_T and total alkalinity were measured
 162 regularly to verify instrument performance (Dickson, 2001). A time-series of CRM measurements over the duration
 163 of the work described here showed no systematic drift. (Fig. A1 in Appendix A). To assess if the change in pH was

164 driven by the addition of CO₂, the final pH and available C_T measurements were compared with a model described
165 here. The theoretical change in tris-artificial seawater (ASW) pH due to an increase in C_T is straightforward to
166 calculate, since both tris and CO₂ acid-base equilibria are well-characterized in seawater and ASW media. The pH is
167 calculated for tris-ASW + C_T using an equilibrium model following the approach described in Chapter 2 of Dickson
168 et al. (2007) for the case of known alkalinity and C_T. In the case of ASW, the seawater equilibrium constants for CO₂
169 are appropriate because minor ions present in seawater and not ASW do not appreciably affect the CO₂ equilibrium
170 constants (particularly when the goal is to compute relative changes in pH) as the ionic background of ASW is closely
171 matched to that of seawater at salinity = 35. In our model, minor acid-base species important to seawater alkalinity
172 but not present in ASW (borate, phosphate, silicate, fluoride) are set to zero. The definition of total alkalinity is
173 modified to include the tris acid-base system following the definition of acid-base donor/acceptor criteria given by
174 Dickson (1981): tris is assigned as a level-1 proton acceptor and tris-H⁺ is at the zero level. Thus, in our model, tris_{tot}
175 = 0.08 molal and alkalinity = 0.04 molal and C_T is a variable. An algorithm (see Annexe 1 in Dickson et al. (2007)) is
176 then used to find the root of the alkalinity equation in its residual form by solving for pH.

177 3. Results & Discussion

178 Figure 2 depicts pH_{spec,20°C}, stored in either a bag or bottle, as a function of time and is subdivided for tests 1,
179 2, and 3. A linear decrease was observed for all bags or bottles. A linear regression was calculated for each
180 experimental condition and, in the cases where measurements at t = 0 were removed due to protocol changes described
181 above, the line was extrapolated back to t = 0, shown by the dotted line. The measured or extrapolated y-intercept is
182 reported as the initial pH in Table 2. In all tests, trendlines are extrapolated to t = 365 days to illustrate observed and

183 predicted change over the course of a year as shown by the solid line. For ease of visual comparison, the y-axis of
 184 each subplot has an identical pH range of 0.017.
 185



186
 187 **Figure 2: Individual time series of measured pH in tris buffer solutions. Tris batch is indicated by shape, storage vessel by**
 188 **color, and storage location by fill. This marker system is also followed in Fig. A2. The solid line is a linear regression starting**
 189 **at the first included pH measurement and ending 365 days after the tris was bagged. The dotted line illustrates the**
 190 **extrapolation back to 0 days stored in bag when measurements at $t = 0$ do not exist. The range of the y-axis scale is fixed at**
 191 **0.017 pH for all subplots.**

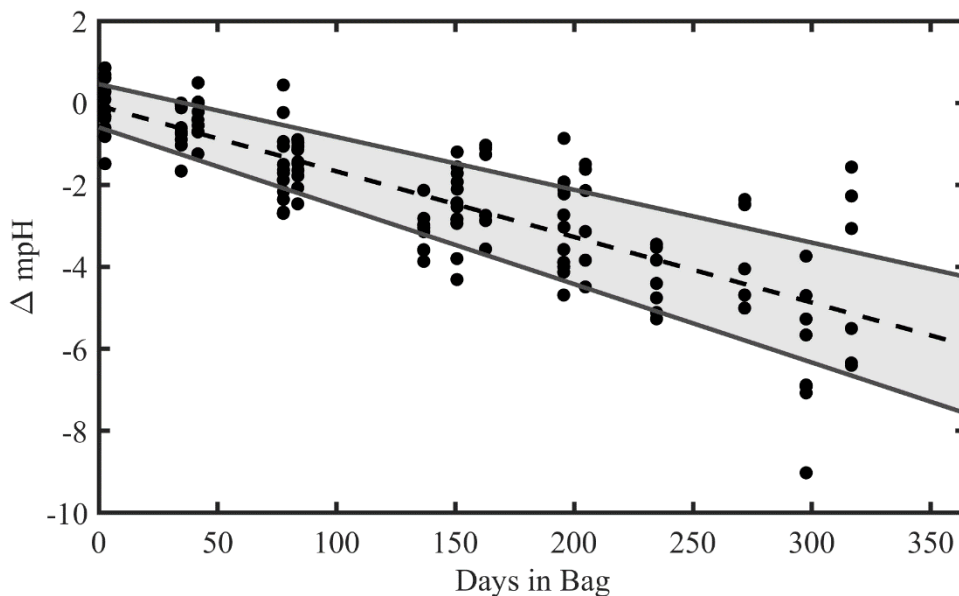
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193 **Table 2: Linear regression statistics from trendlines shown in Fig. 1 and 2. The last row shows the regression statistics for**
 194 **tris from all batches, in either bag type, stored in the lab or test tank. Slope and intercept are shown as mean \pm 95%**
 195 **confidence intervals. The reported intercept is the regression intercept; when initial pH measurements are available, they**
 196 **differ by less than 0.0003 from regression intercept. * Indicates the outlier (Batch 2, Bag 1, Lab) caused by a damaged bag.**
 197 **The outlier, “Batch 2, Bag 1, lab”, was not used in the “All Batches, All Bags, Lab or Tank” composite. † In all batches, all**
 198 **bags, lab or tank, the slope was calculated with a linear fit of all (non-outlier) tris measurements. The RMSE is the mean**
 199 **RMSE of all (non-outlier) bag fits. ‡ The calculated tris pH was calculated at 20°C; however, this calculated pH is 0.0135**
 200 **higher than equimolar tris as noted above (DeValls and Dickson, 1998).**

Batch & Storage Method	Slope (mpH yr ⁻¹)	Intercept (Initial pH)	RMSE (mpH)	r ²	n
Batch 1, Bag 1, Lab	-2.9 \pm 1.7	8.2558 \pm 0.0012	0.43	0.59	12
Batch 2, Bag 1, Lab*	-13.7 \pm 2.7	8.2627 \pm 0.0018	0.61	0.94	11
Batch 2, Bag 2, Lab	-6.7 \pm 2.2	8.2519 \pm 0.0015	0.55	0.82	12
Batch 3, Bag 1, Lab	-4.6 \pm 2.7	8.2539 \pm 0.0010	0.62	0.62	11
Batch 3, Bag 1, Tank	-6.9 \pm 3.2	8.2683 \pm 0.0012	0.73	0.73	11
Batch 3, Bag 2, Lab	-3.8 \pm 2.1	8.2462 \pm 0.0008	0.54	0.61	12
Batch 3, Bag 2, Tank	-7.9 \pm 2.1	8.2335 \pm 0.0008	0.44	0.92	9
Batch 4, Bag 1A, Lab	-6.4 \pm 1.3	8.2630 \pm 0.0005	0.64	0.90	14
Batch 4, Bag 1B, Lab	-5.8 \pm 1.8	8.2631 \pm 0.0008	0.91	0.79	15
Batch 4, Bag 1C, Lab	-8.0 \pm 1.0	8.2631 \pm 0.0004	0.49	0.96	15
Batch 4, Bag 1D, Lab	-9.2 \pm 1.6	8.2631 \pm 0.0007	0.80	0.92	15
Batch 4, Bottle, Lab	-3.0 \pm 1.4	8.2638 \pm 0.0005	0.81	0.44	25
All Batches, All Bags, Lab or Tank†	-5.8 \pm 1.1	–	0.72	0.66	126
Calculated tris pH‡	–	8.2652	–	–	–

201
 202 Only bags from test 3, using tris batch 4 and bag type 1, have direct initial pH measurements and replicate
 203 bags. Initial pH measurements of these 4 bags were 8.2630 \pm 0.0007 (mean \pm standard deviation, n = 12). Importantly,
 204 the very low standard deviation suggests that a single initial pH measurement is representative of all replicate bags
 205 filled with a single tris batch, if the preparation procedure used in test 3 is followed. This inter-bag consistency is
 206 beneficial because it reduces the number of initial pH measurements required when filling multiple bags. There is also
 207 strong agreement in initial pH measurements between bagged and bottled tris in test 3, with the initial pH of bottled
 208 tris 0.0007 higher than bagged tris (8.26327 \pm 0.0004, n = 6). The differences in filling procedure or impurities between
 209 bags and bottles in test 3 appear to have little effect on the initial pH. The mean initial pH of tris batch 4 is 0.002 (n =
 210 5) lower than calculated pH_{tris,20°C} (Fig. A2). This difference between the mean initial pH of tris batch 4 and calculated
 211 pH_{tris,20°C} is similar in direction and magnitude to those reported in other studies: DeGrandpre et al. (2014) reported –
 212 0.0012 \pm 0.0025 and Müller and Rehder (2018) reported -0.002 to -0.008 (measured pH minus pH_{tris,T_C}). With standard
 213 laboratory equipment and off-the-shelf reagents, an uncertainty of 0.006 is expected in prepared tris (Paulsen and
 214 Dickson, 2020). Measurements were also made on Dickson standard tris (batch T35) using the same instrument and
 215 the pH was 0.0019 higher than the calculated pH_{tris,20°C} (n = 2). In tests 1 and 2, the initial pH was extrapolated from
 216 a linear regression. The extrapolated initial pH values are more variable and lower (on average) than those directly

217 measured (Fig. A2). These differences may be a result of the extrapolation or different experimental variables such as
 218 the increased rinsing of bags, or the single bag type and storage location used in test 3.

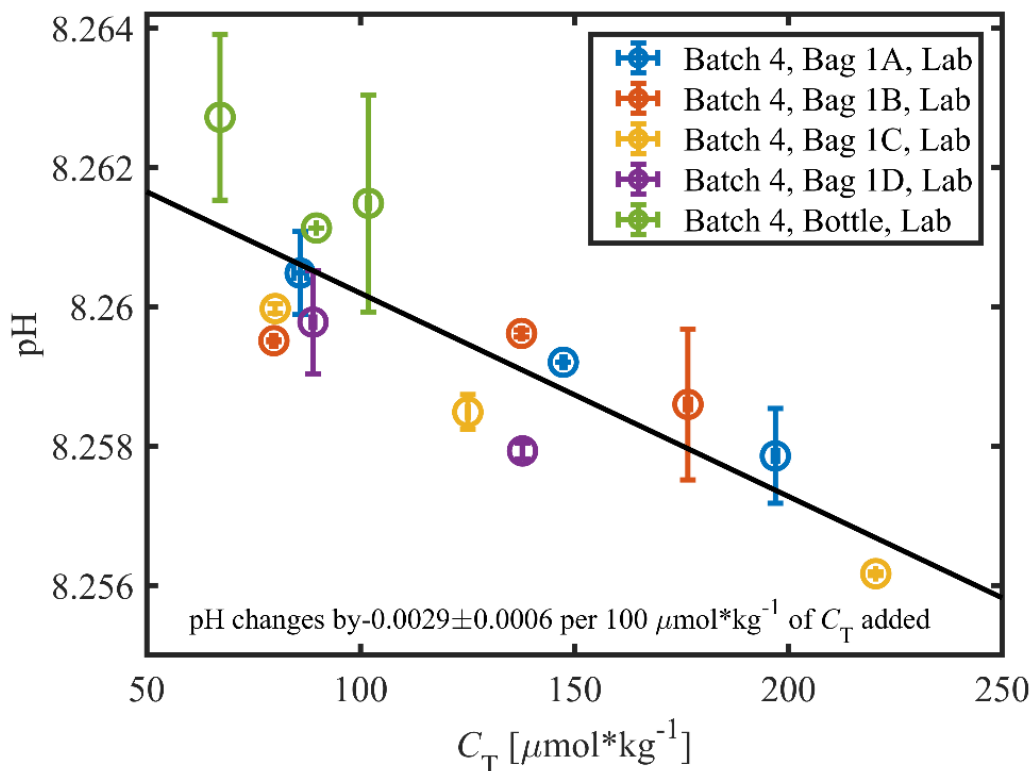


219
 220 **Figure 3: Combined time series of measured pH in tris buffer tris buffer solutions. The dots represent every**
 221 **measurement made on a (non-damaged) bag of tris. The dotted line is the “All Bags, All Batches, Lab or Tank”**
 222 **regression. The grey shaded region is the observational 95% confidence interval (CI). The CI is intended to**
 223 **estimate the future pH of a tris bag (with known initial pH and an unmeasured bag specific rate of change).**
 224 **The upper and lower bounds are -0.0042 and -0.0076 pH per year, respectively.**

225 Figure 3 depicts a composite of all test results as the change from the initial pH of tris
 226 ($\Delta pH = pH_{spec,20^{\circ}C}^{t=day} - pH_{spec,20^{\circ}C}^{t=0}$) as a function of time elapsed since bagging. A linear regression on all pH
 227 measurements, excluding the outlier of “Batch 2, Bag 1, Lab”, of tris stored in bag types 1 or 2, has a slope of -0.0058
 228 $\pm 0.0011 \text{ yr}^{-1}$ (mean \pm 95% C.I.). The upper and lower bounds of ΔpH at $t = 365$ days, -0.0042 and -0.0076 , are
 229 important to consider when utilizing this bagged storage method of tris. These bounds provide the broadest expected
 230 range in pH change over a year of storage and include both the intercept and slope confidence intervals ($slope_{CI}$ and
 231 $intercept_{CI}$, respectively). For example, the upper bound of ΔpH at $t = 365$ days is calculated as: $upper\ bound =$
 232 $(slope + slope_{CI}) * 365 + intercept + intercept_{CI}$. The outlier (Batch 2, Bag 1, Lab) was excluded due to
 233 noticeable damage to the bag (see Fig. A3 in Appendix A), which is believed to have caused its pH to decrease at
 234 more than two times the average rate of the other bags. The damage appears to be a break in the metallic bag layer,
 235 potentially caused by creasing or pinching of the bag when handling. This observation highlights the importance of
 236 maintaining bag integrity, particularly during use in the field. A successful two-week field deployment has been
 237 conducted using the tris bags described here and a modified SeapHOx in a shallow, coral reef flat (Bresnahan et al.
 238 2021). This two-week deployment was significantly shorter than the year of storage described here and further field
 239 testing in longer deployments in varied environments are required before widespread use of this technology. For the
 240 longer time frame depicted in Figure 3, the only comparable example found in the literature is the work of Lai et al.

241 (2018). In this work, Lai et al. (2018) used bagged tris for sensor calibration, with in situ tris measurements made over
 242 150 days. Lai et al. (2018) did not report a change in the pH of bagged tris over the deployment; however, the reported
 243 precision of the SAMI-pH in situ instrument (± 0.003) would not resolve the expected change shown in our Figure 3.
 244 Therefore, the results of Lai et al. (2018) are not inconsistent with our study.

245 A significant increase in C_T was observed for all types of bags and bottles in Experiment 3 (Figure 4). A high
 246 correlation between solution pH and C_T was observed, with a slope of -0.0029 ± 0.0006 pH per $100 \mu\text{mol kg}^{-1}$ ($n =$
 247 14 , $r^2 = 0.70$), suggesting that the change in tris pH and C_T was primarily driven by an increase in CO_2 . The observed
 248 slope agrees closely with a theoretical model prediction of a linear decrease in pH of -0.0024 per $100 \mu\text{mol kg}^{-1}$ of
 249 C_T added (over the range of C_T observed). There are two possible sources of the increasing C_T : gas exchange of CO_2
 250 with the environment and microbial respiration within the storage vessel. Gas exchange should not be a significant
 251 source of CO_2 for tris stored in a borosilicate bottle, as this is the standard equipment used to store seawater CO_2 and
 252 tris buffers and is known to minimize gas exchange (Dickson et al. 2007). Therefore, it is likely that respiration was
 253 the primary driver for the increase in C_T for tris stored in bottles. On average, pH decrease of tris stored in bags was
 254 larger than that in the standard bottle (Figure 2), indicating either an additional source of CO_2 from gas exchange, or
 255 larger amounts of respiration. Distinguishing between these two theorized sources would require measurements of
 256 additional parameters such as dissolved organic carbon.



257
 258 **Figure 4: pH plotted against C_T shows a linear relationship between the two parameters in tris buffer with a slope of $-$**
 259 **0.0029 pH for every $100 \mu\text{mol kg}^{-1}$ of C_T added. The measurements shown are from three sampling occurrences between**
 260 **130–300 days stored on bags and bottles used in Test 3. Only two measurements are shown for “Batch 4, Bag 1D, Lab”**
 261 **because it ran empty before C_T were made.**

262 The pH stability of tris could be improved by reducing either likely source of C_T : gas exchange or microbial
263 respiration. For bags, CO_2 may diffuse through the fittings, gasket, or bag walls, particularly if damaged. The relatively
264 small breaks in the aluminium foil layer caused “Batch 2, Bag 1, Lab” to decrease more than twice as fast as the
265 average bag. Storage bag, fitting, and gasket material, as well as careful handling, are therefore important factors in
266 minimizing gas exchange. For example, silicone is permeable to CO_2 , and thus could have been a path of gas exchange
267 into the tris for this experiment. As noted above, Nemzer and Dickson (2005) found an almost negligible change of
268 0.5 mpH yr^{-1} in bottled tris. Our bottled tris changed at -3.0 mpH yr^{-1} ($n = 10$ bottles measured over 161 days),
269 approximately half the rate of the tris stored in bags. While -3.0 mpH yr^{-1} is near the detection limit of our
270 measurements, it suggests that the bottling protocol used in this study was not as well controlled as that of Nemzer
271 and Dickson (2005). For example, the Dickson lab at Scripps Institution of Oceanography regularly uses an annealing
272 oven to combust all trace organic films that may persist on glass bottles, but in our study, bottles were not annealed.
273 Although bags cannot be annealed, future steps that may be worth consideration to reduce microbial respiration in
274 bags include addition of a biocide to the tris solution, acid cleaning the bags, and using ultraviolet light to remove
275 organics from the ultrapure water used to prepare tris. There are some disadvantages to these proposed steps. Addition
276 of a biocide may not be ideal for use in sensitive environments if the tris is discharged after use and would alter the
277 composition of the solution slightly. While rinsing or prolonged soaking of the bags with an acid may help to remove
278 organics, it is unclear if it would have negative effects on the integrity of the bags. Beyond removing organics on the
279 bag surfaces, care should be taken to avoid introducing organic contaminants into the tris during the solution
280 preparation and bag filling procedures to minimize future respiration.

281 Both bag type 1 and 2 experienced problems with structural integrity during this experiment. A single type 2
282 bag experienced delamination of exterior bag layers when stored submerged in seawater, causing the eventual tearing
283 and failure of the bag when handling. Bag type 2 was not used in test 3 due to this failure. It should be noted that in
284 other studies which successfully used bag type 2, the bag was submerged in seawater for less time than in this
285 experiment (Sayles and Eck, 2009; Aßmann et al., 2011; Wang et al., 2015). A single bag type 1 had the subtler
286 problem of small breaks in the aluminium foil bag layer, likely causing an increased pH rate of change. In non-
287 damaged bags, factors such as bag type/bottle, lab/tank storage, or tris batch did not have statistically significant (p -
288 value < 0.05) correlations with the pH change of tris (p -values 0.12, 0.11 and 0.09, respectively). The results of the
289 ANOVA support that tris can be held in bag type 1 or 2 and stored in a lab or tank and the pH will change similarly
290 regardless of storage method for up to 300 days. Additional bag types could be tested, such as bags made by Pollution
291 Measurement Corp. used by Lai et al. (2018) or Scholle DuraShield used by Takeshita et al. (2015).

292 These results suggest that when bags are carefully handled prior to and after filling, tris pH changes are small
293 over time. Specific recommendations for further work include: bags must be handled with care and enclosed in
294 protective containers to prevent damage, bags must be rinsed with tris prior to filling, and additional testing is merited
295 to determine sources of and methods to reduce contamination, such as acid washing.

296 **4. Conclusions**

297 This article describes our characterization of the stability of tris buffer in artificial seawater when stored in
298 purportedly gas-impermeable bags. Several different tests, initiated over the course of a year and a half and lasting up
299 to 300 days, exhibited an average decrease of 5.8 mpH yr⁻¹. In comparison, tris stored in standard borosilicate bottles
300 was shown to have a decrease of 3.0 mpH yr⁻¹. For yearlong deployments, an expected pH change of -0.0058 is well
301 below the weather quality threshold of 0.02 pH units. This low rate of change demonstrates the value of bagged tris
302 for in situ validation of autonomous pH sensors (regardless of sensor operating principles), particularly in highly
303 dynamic areas where repeatability of calibration based on discrete samples is challenging. Given the thorough
304 characterization of tris over wide ranges of environmental variables, this contribution can aid in the traceability and
305 intercomparability of pH sensor measurements. While valuable at the current stage of development (as demonstrated
306 by, e.g., Lai et al. (2018) and Bresnahan et al. (2021)), further development would ideally result in a commercially
307 available bag and filling procedure that can yield a rate of pH change less than the climate threshold of 0.003 per year.
308 This will require further tests to identify the source of CO₂, gas exchange or microbial respiration, as well as steps to
309 reduce or eliminate these sources.

310 Periodic measurement of bagged tris in situ would allow for detection of sensor drift. Most in situ pH sensors
311 are deployed in the euphotic zone in coastal areas, typically resulting in expedited biofouling and sedimentation, and
312 leading to sensor drift (Bresnahan et al., 2014) that could be identified and potentially corrected. Such periodic
313 calibration/validation would aid in identifying sensor issues and allow for greater consistency and continuity between
314 a timeseries and planned or vicarious crossovers where an automated calibration can be used to augment or replace
315 pre- and post-deployment calibrations/validations.

316 **Appendix A**

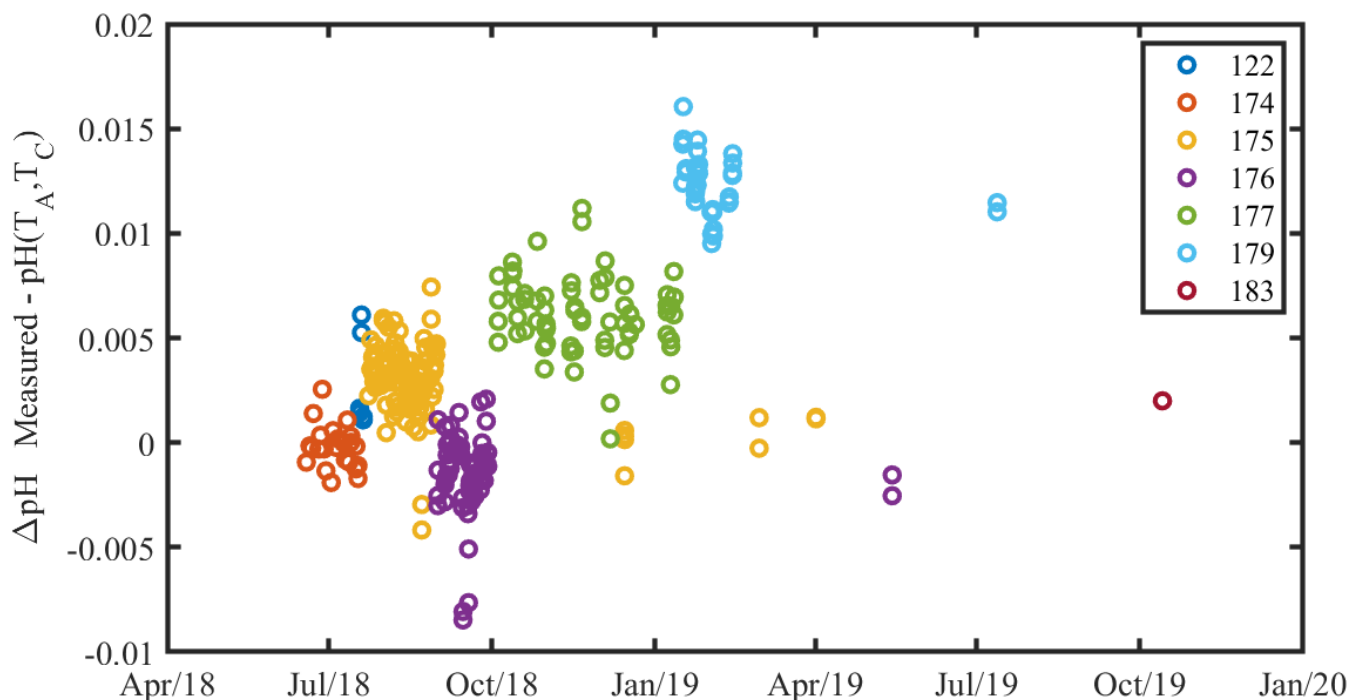
317

318 **Table A1. Detailed information about the specific reagents used to make the tris solution. *Reagent chemicals that meet or**
 319 **surpass specifications of the British Pharmacopeia (BP), European Pharmacopeia (EP), Food Chemicals Codex (FCC),**
 320 **United States Pharmacopeia (USP).**

321

Chemical	Manufacture	Part Number	Lot Number	Batch	Assay	Grade
tris	Fisher Scientific	T395-1	170360	all	99.8%	Certified ACS
NaCl	Fisher Scientific	S641-212	127252	all	99.0 to 100.5%	*BP/EP/FCC/USP
Na ₂ SO ₄	Fisher Scientific	S421-1	134837	all	99.8%	Certified ACS
KCl	Fisher Scientific	P217-500	174416	all	99.7%	Certified ACS
MgCl ₂	Teknova	M0304	M030427E1401	all	1 M	Biotechnology
CaCl ₂	Amresco	E506-500mL	0982C098	all	0.95-1.05 M	Biotechnology
HCl	Fisher Scientific	SA48-1	175004	1, 2, 3	0.999 N	Certified
HCl	Fisher Scientific	SA48-1	188768	4	1.003 N	Certified

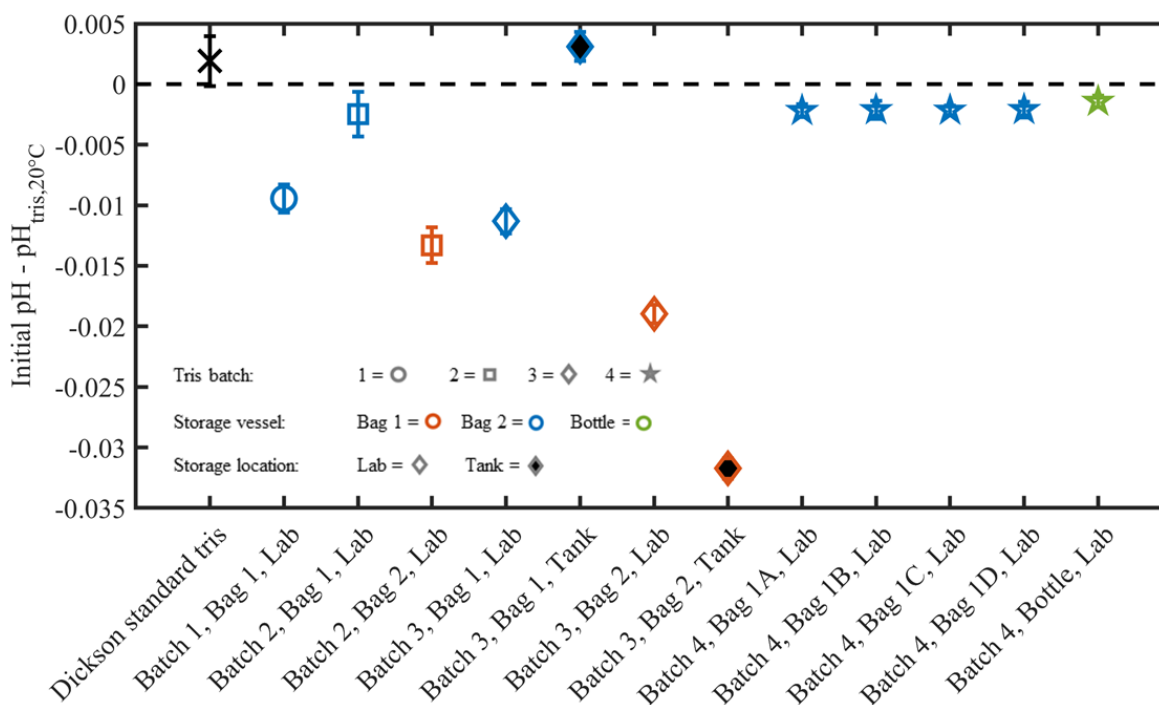
322



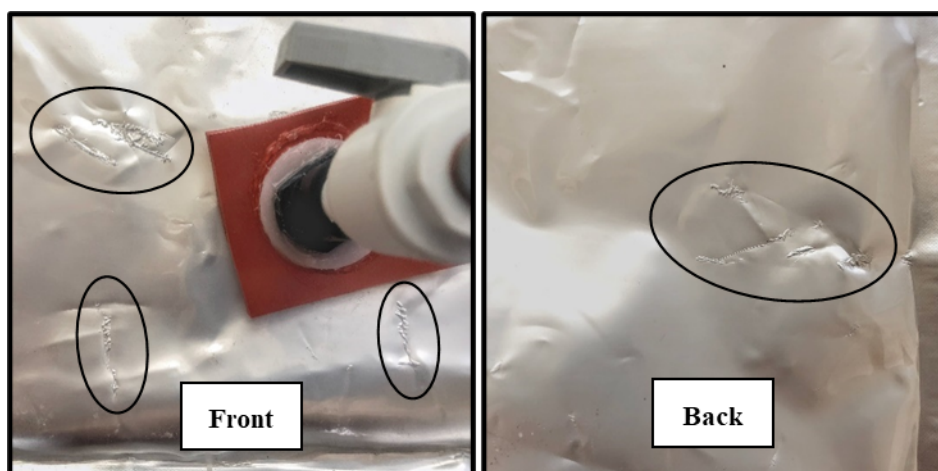
323

324 **Fig. A1: A timeseries of the residual between measured and calculated CRM pH throughout the experiment. Marker color**
 325 **denotes CRM batch number. There is a clear variability between measured and calculated pH, which typical of CRM**
 326 **batches (Andrew Dickson, *pers. comm.*). There was no observable systematic drift in the pH system during the experiment.**
 327 **The mean standard deviation of pH measurements within a CRM batch is 0.0016, which is comparable to the 0.0019**
 328 **reported in Bockmon & Dickson (2015). The same 760 nm absorbance wavelength outlier removal procedure used for tris**
 329 **measurements was applied to CRM measurements.**

330



331
 332 **Fig. A2: The initial pH residual of each tris bag or bottle measured in this experiment. The initial pH is reported as a**
 333 **residual from the calculated pH at 20 °C. The initial pH was measured directly for tris batch 4 and extrapolated for tris**
 334 **batches 1-3. Additionally, 2 bottles of Dickson standard tris (show by the black “X”) were measured on 12/10/2018. The**
 335 **zero black dashed line is the calculated pH of tris at 20 °C, based upon the measured reagent concentrations (DelValls and**
 336 **Dickson, 1998).**



337
 338 **Fig. A3. The ovals indicate marks on the exterior of “Batch 2, Bag 1, Lab”. These marks appear to be damage to the interior**
 339 **metallic layer, possibly due to creasing of the bag. These marks were not present on any other bag used in this study.**

340
 341

342 **Author contribution**

343 WW performed formal analysis, visualization, and writing – original draft preparation. KS and TW contributed
344 to investigation and writing – review & editing. PB, YT, and TM contributed to funding acquisition, conceptualization,
345 formal analysis, and writing – review & editing.

346 **Competing interests**

347 The authors declare that they have no conflict of interest.

348 **Data availability**

349 pH and C_T data are available via the UC San Diego Library Digital Collections at
350 <https://doi.org/10.6075/J0QC022G> (Wolfe et al., 2021).

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