Ocean Sci. Discuss., 6, S154–S155, 2009 www.ocean-sci-discuss.net/6/S154/2009/
© Author(s) 2009. This work is distributed under the Creative Commons Attribute 3.0 License.



OSD

6, S154-S155, 2009

Interactive Comment

## Interactive comment on "Salinity/temperature ranges for application of seawater $S_A$ -T-P models" by G. M. Marion et al.

G. M. Marion et al.

Received and published: 9 June 2009

Here are my comments with respect to Reviewer #3 arranged by his/her numbers. 1. I changed the title to "Precipitation of Solid Phase Calcium Carbonates and Their Effect on Application of Seawater SA-T-P Models." 2. Eqs. 1-4 do not deal with the precipitation of pure CaCO3(calcite). The equilibrium constant K(calcite) is for pure calcite, but only as a reference point. But the IAP(CaCO3) is based on experimental measurements of seawater CaCO3 that likely contained Mg2+ and could have been either aragonite or calcite. The omega(calcite) refers to calcite only because the reference state is pure calcite. These formulations could also have been represented as omega(aragonite) by dividing IAP(CaCO3) by the pure aragonite equilibrium constant. In the revision, I rewrote Eq. 3 to show the calcite-aragonite relationship. 3. I rewrote the paper to present MPa as the dominant pressure term, except for atmospheric val-

Full Screen / Esc

Printer-friendly Version

Interactive Discussion

**Discussion Paper** 



ues that I prefer to represent as μatm. But whenever different terms for pressure are given in the text, I added parentheses that list the alternative terms. For example, 0.1-100 MPa (= 1 to 1000 bars). 4a. PCO2 was kept constant at 385 μatm. The listed values in Table 1 can only lead to minor changes as T or PCO2 change. Nevertheless, while alkalinity remains fairly constant, the components of alkalinity (now defined in a footnote in Table 1) can significantly change. So this can cause a small change in salinity, but this is minor (≈ 0.04%). This point is now described in more detail in Section 2.2 (Seawater Properties). 4b. The only strong point in support of these extrapolations is the fact that the model accurately predicts the transition from calcite to ikaite at 3°C, which is in excellent agreement with the literature. I embellished my discussion on this subject in Section 4. 5. I rewrote the line as suggested by the reviewer. 6. 450 μatm is correct. Gypsum starts precipitating sometime between 2008 and 2100, approximately when the PCO2 reaches 450 μatm. I changed the wording to clarify. 7. I reworded as suggested by the reviewer. 8. I added PCO2 to Figure 2. I now mention in Section 2.2 how I deal with PCO2 and fCO2 in this paper.

Interactive comment on Ocean Sci. Discuss., 6, 171, 2009.

## OSD

6, S154-S155, 2009

Interactive Comment

Full Screen / Esc

Printer-friendly Version

Interactive Discussion

Discussion Paper

